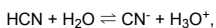


HW05 - Acids, Bases, and Salts

Question 1

1 pts

In the reversible reaction



the two Bronsted-Lowry acids are...

- There is only one Bronsted-Lowry acid shown: H_3O^+ .
- HCN and H_3O^+
- H_2O and H_3O^+
- HCN and CN^-
- H_2O and CN^-

Question 2

1 pts

A water solution of sodium acetate is basic because...

- sodium acetate is only weakly ionized.
- The statement is false. A water solution of sodium acetate is acidic.
- the conjugate base of the acetate ion is a strong base.
- the acetate ion acts as a Bronsted-Lowry base in a reaction with water.

Question 3

1 pts

According to the Bronsted-Lowry concept of acids and bases, which of the following statements about a base is NOT true?

- If a base is strong, then its conjugate acid will be relatively weaker.
- A base will share one of its electron pairs to bind H^+ .
- A base reacts with an acid to form a salt.
- A base must contain a hydroxide group.

Question 4

1 pts

Which of the following is true in pure water at any temperature?

- $[\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$
- $[\text{H}_3\text{O}^+] = [\text{OH}^-]$
- K_w decreases with increasing temperature.
- $\text{pH} = 7.0$

Question 5

1 pts

What is $[\text{H}_3\text{O}^+]$ when $[\text{OH}^-] = 3.3 \times 10^{-9} \text{ M}$?

- $3.0 \times 10^{-6} \text{ M}$
- $3.3 \times 10^{-9} \text{ M}$
- $3.3 \times 10^{-5} \text{ M}$
- $1.0 \times 10^{-7} \text{ M}$

Question 6

1 pts

A strong acid (or base) is one which...

- should only be used when wearing goggles and gloves.
- reacts with a salt to form water.
- dissolves metals.
- dissociates completely in aqueous solution.

Question 7

1 pts

Which of the following substances is a strong acid?

- H_2SO_4
- H_3PO_4
- HF
- HSO_3
- H_2CO_3

Question 8

1 pts

HCN is classified as a weak acid in water. This means that it produces...

- no hydronium ions.
- a relatively large fraction of the maximum number of possible hydronium ions.
- a relatively small fraction of the maximum number of possible hydronium ions.
- 100% of the maximum number of possible hydronium ions.

Question 9

1 pts

Which of the following substances is a weak acid?

- HNO₃
- HI
- HClO₄
- H₂SO₄
- HCl
- HClO₃
- HBr
- H₂CO₃

Question 10

1 pts

Which is NOT a conjugate acid-base pair, respectively?

- H₂O : OH⁻
- SO₄²⁻ : HSO₄⁻
- HCN : CN⁻
- H₃O⁺ : H₂O

Question 11

1 pts

The conjugate base of H₂SO₄ is:

- HSO₄⁻
- HSO₄
- SO₄²⁻
- H₃SO₄⁺

Question 12

1 pts

What is the conjugate acid of NO₃⁻?

- NO₃²⁻
- NH₃
- HNO₃
- NO₂⁻

Question 13

1 pts

Assume that five weak acids, identified only by numbers (1, 2, 3, 4, and 5) have the following ionization constants:

- 1 - 1.0×10^{-3}
- 2 - 3.0×10^{-5}
- 3 - 2.6×10^{-7}
- 4 - 4.0×10^{-9}
- 5 - 7.3×10^{-11}

The anion of which acid is the strongest base?

- 3
- 4
- 2
- 5
- 1

Question 14

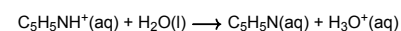
1 pts

The term "K_a for the ammonium ion" describes the equilibrium constant for which of the following reactions?

- NH₄⁺ + OH⁻ ⇌ NH₃ + H₂O
- NH₃ + H₂O ⇌ NH₄⁺ + OH⁻
- NH₄Cl(solid) + H₂O ⇌ NH₄⁺ + Cl⁻
- NH₄⁺ + H₂O ⇌ NH₃ + H₃O⁺

Question 15

1 pts

If the value of K_b for pyridine (C₅H₅N) is 1.8×10^{-9} , calculate the equilibrium constant for the following reaction:

- -1.8×10^{-9}
- 1.8×10^{-16}
- 5.6×10^{-6}
- 5.6×10^8

Question 16

1 pts

What is $[\text{OH}^-]$ in a 0.0050 M HCl solution?

- 1.0×10^{-7} M
- 6.6×10^{-5}
- 2.0×10^{-12} M
- 1.0 M

Question 17

1 pts

Which pH represents a solution with 1000 times higher $[\text{OH}^-]$ than a solution with a pH of 5?

- pH = 4
- pH = 6
- pH = 8
- pH = 7

Question 18

1 pts

What is the pH of a 0.1 M $\text{Ba}(\text{OH})_2$ aqueous solution?

- 1.33
- 13.3
- 9.98
- 8.7

Question 19

1 pts

Hydroxylamine is a weak molecular base with $K_b = 6.6 \times 10^{-9}$. What is the pH of a 0.0500 M solution of hydroxylamine?

- 8.93
- 10.4
- 9.48
- 9.26

Question 20

1 pts

What is the pH of a 0.23 M solution of potassium generate (KR-COO)? K_a for the generic acid R-COOH is 2.7×10^{-8} .

- 10.23
- 10.47
- 10.83
- 10.60

Question 21

1 pts

Which solution has the highest pH?

- 0.1 M KClO , K_a for HClO is 3.5×10^{-8}
- 0.1 M KCH_3COO , K_a for CH_3COOH is 1.8×10^{-5}
- 0.1 M of KNO_2 , K_a for HNO_2 is 4.5×10^{-4}
- 0.1 M of KCl , K_a for HCl is VERY LARGE!!

Question 22

1 pts

What is the pH of a solution that contains 11.7g of NaCl for every 200 mL of solution?

- 1.0×10^{-7}
- 9.0
- 10^{-1}
- 7.0

Question 23

1 pts

What is the pH of a solution made by mixing 0.050 mol of NaCN with enough water to make a liter of solution? K_a for HCN is 4.9×10^{-10} .

- 12
- 11
- 10^{-3}
- 3

Question 24

1 pts

Identify the list in which all salts produce a basic aqueous solution.

- NH_4Cl , $\text{C}_6\text{H}_4\text{NH}_3\text{NO}_3$, FeI_3
- AlCl_3 , $\text{Zn}(\text{NO}_3)_2$, KClO_4
- KCH_3COO , NaCN , KF
- AgNO_3 , NaCHO_2 , CrI_3

Question 28

1 pts

A 0.28 M solution of a weak acid is 3.5% ionized. What is the pH of the solution?

- 1.46
- 2.01
- 3.17
- 0.55

Question 25

1 pts

What is the pH in a solution made by dissolving 0.100 moles of sodium acetate (NaCH_3COO) in enough water to make one liter of solution? K_a for CH_3COOH is 1.80×10^{-5} .

- 10.25
- 9.25
- 8.87
- 5.74

Question 29

2 pts

The pH of 0.010 M aqueous aniline is 8.32. What is the percentage protonated?

- 0.021%
- 2.1%
- It is impossible to tell without knowing the K_a or the K_b for aniline.
- 0.0021%

Question 26

1 pts

A 0.200 M solution of a weak monoprotic acid HA is found to have a pH of 3.00 at room temperature. What is the ionization constant of this acid?

- 5.0×10^{-6}
- 1.0×10^{-3}
- 2.0×10^{-9}
- 5.3

Question 27

1 pts

What is the percent ionization for a weak acid HX that is 0.40 M? $K_a = 4.0 \times 10^{-7}$.

- 0.0010%
- 0.10%
- 0.0020%
- 0.20%